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Exercise Series N° 1

Exercise 1:

1) List the different units of pressure and provide the relationships between these units.
2) Convert 350 mmHg to bar, Pa, and atm.
3) State the different temperature scales and provide the relationships between these scales.
4) Convert: 25°C and -100°C to °F, 0°F to °C and K, 0 K to °C and °F.
5) Calculate the value of the ideal gas constant R in $L.atm.mol^{-1}.K^{-1}$, $J.mol^{-1}.K^{-1}$, and $cal.mol^{-1}.K^{-1}$, knowing that one mole of an ideal gas occupies a volume of 22.4 L at a pressure of 1 atm and a temperature of 0°C.

Exercise 2:

An ideal gas initially in an equilibrium state characterized by P_1=2 atm and V_1=2L undergoes an isothermal expansion until P_2=0.5 atm. What is the volume of the final state?

Exercise 3:

A mass of nitrogen (assumed to be an ideal gas) occupies a volume of 20L at 20°C. It is heated under constant pressure (isobaric process). What volume does it occupy at 120°C?

Exercise 4:

A quantity of air, assumed to be an ideal gas (M=29g/mol), is under a pressure of 10 bars and a temperature of 5°C, contained in a closed and perfectly rigid chamber with a capacity of 5L. Determine the gas pressure at 80°C and the volumetric mass in both equilibrium states.

Exercise 5:

The analysis of a 100g sample of air collected at sea level yields the following results: N_2: 75.52%, O_2: 23.15%, Ar: 1.28%, CO_2: 0.046%.

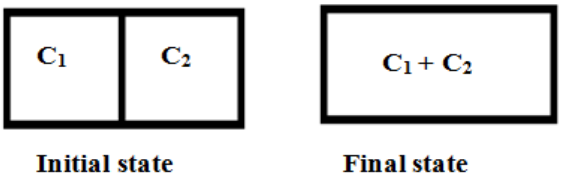
1) Calculate the number of moles of each gas present in this sample.
2) Calculate the mole fraction as well as the partial pressure of each gas.

Data: P_t = 1 atm. N(14), O(16), Ar(40), C(12).

Exercise 6:

A closed, perfectly rigid horizontal cylinder is divided into two compartments, C_1 and C_2, by a partition. At the initial state, compartment C_1 contains 10 moles of oxygen (O_2) at temperature T_1=27°C and pressure P_1=2 atm, and compartment C_2 contains 20 moles of nitrogen (N_2) at temperature T_2=127°C and pressure P_2=3 atm. The partition is removed, and the equilibrium temperature of the mixture T_m reaches 90°C. Assuming both gases are ideal, determine:

1) The pressure of the mixture.
2) The partial pressure of O_2 and N_2.
3) The volumetric mass of the mixture.
4) The density of the mixture.



Data: N: 14g/mol, O: 16g/mol